

## Chemistry 13 Group Problem

Calculate the pH of the following solutions:

1. 0.10 M  $\text{HClO}_3$  (chloric acid)
2. 0.10 M  $\text{HC}_3\text{H}_3\text{O}_3$  (pyruvic acid)
3. 0.10 M  $\text{H}_3\text{AsO}_4$  (arsenic acid)
4. 0.10 M  $\text{NaNO}_3$  (sodium nitrate)
5. 0.10 M  $\text{KCNO}$  (potassium cyanate)
6. 0.10 M  $\text{Na}_2\text{CO}_3$  (sodium carbonate)
7. 0.10 M  $\text{K}_2\text{O}$  (potassium oxide)
8. 0.10 M  $\text{C}_5\text{H}_5\text{N}$  (pyridine)
9. 0.10 M  $(\text{CH}_3\text{NH}_3)(\text{NO}_3)$  (methyl**ammonium** nitrate)