many ionic’s water solubility; most molecular (other solvents), ↑ with \( T \)

Structure: Solute/Solvent

ion-H-bonding, ion-dipole, H-bonding, dipole-dipole

Hydrophobic?
oil often called hydrophobic, implying oil repels water
“water fearing”

\( \Delta H_1 \) Not true! \( \Delta H_1 \)

\( H_2O \)oil dipole-London stronger than oil’s London force;
but not strong enough to break water H-bonds!

Phase Diagram of an Aqueous Solution

Osmotic Pressure

reverse osmosis
water desalination plant